

AQA – Energetics – AS Chemistry P2

1. No.10

Which statement explains why the rate of a reaction increases when a catalyst is added to a reaction mixture at a constant temperature?

[1 mark]

- A** The collision frequency increases because there is an increase in the activation energy.
- B** The collision frequency increases because there is an increase in the mean energy of the particles.
- C** The proportion of successful collisions increases because there is a decrease in the activation energy.
- D** The proportion of successful collisions increases because there is an increase in the mean energy of the particles.

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2. No.19

What is the enthalpy of combustion, in kJ mol^{-1} , of propane?

[1 mark]

A -211.7☐**B** -419.7☐**C** -2220☐**D** -2878☐

3. No.20

What is the enthalpy change, in kJ mol^{-1} , when butane reacts to form propane, hydrogen and carbon?

[1 mark]

A +21.2☐**B** +17.5☐**C** -17.5☐**D** -21.2☐