

AQA – Properties of Period 3 elements and their oxides – A2 Chemistry P3**1. No.19**

Consider the Period 3 elements from sodium to chlorine.

Which statement is correct?

[1 mark]

A Sodium has the smallest atomic radius.

☐

B Aluminium has the highest melting point.

☐

C Sulfur is the most electronegative.

☐

D Chlorine has the highest first ionisation energy.

☐**2. No.20**

Which statement correctly describes a trend down Group 7 from Cl to I?

X represents Cl, Br or I

[1 mark]

A The boiling point of HX increases.

☐

B The bond dissociation energy of H–X increases.

☐

C The standard electrode potential value for $X_2(aq) + 2e^- \rightarrow 2X^-(aq)$ becomes more positive.

☐

D The solubility of AgX in ammonia increases.

☐